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Answer –  $\text{O}_2 \text{ (aq)} + 2 \text{O} \text{ (l)} + 4\text{e}^-$   
 $\text{O}_2 \text{ (aq)} + 2\text{e}^- \rightarrow \text{O} \text{ (g)}$  z ollysis of dilute NaCl  
concentrated NaCl and molten NaCl ent.  $\text{O}_2$  >  $\text{O}$  oduct at

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the anode. (ii) oxygen gas oxygen gas y carbon anode chloride  
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Review Sheet: Unit 6 Name\_\_\_\_\_ CHEMISTRY: A Study of Matter © 2004, GPB 6.38a KEY I. Fill in the blanks with the most appropriate term:

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## ~~Chemistry A Study Of Matter 6 23 Worksheet Balancing ...~~

The error in print is on Page 271 under the Chapter of “  
Electrolysis ” in Chemistry Matters for Pure Chemistry  
syllabus. In the Key Ideas which acts like sub-section  
summary, it states that “ OH- ions will be discharged as  
oxygen gas. However, if a solution is concentrated in halide  
ions, the halogen gas will be produced.”

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~~Error in Marshall Cavendish GCE O Level Chemistry Matters~~

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Assessment Answer Key Chemistry: Matter and Change  
T167 Name Date Class 20 Understanding Main Ideas  
Chemistry: Matter and Change • Chapter 4 Chapter  
Assessment (Part A) Use the periodic table to identify each  
element described below. 1. atomic number 65 2. 78 protons  
3. 44 protons and 44 ...

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