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~~K_w pH pOH pKa~~

~~pKb H^+ OH^-~~

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& Base Test.

Matching: L1.

strong acids A.

bitter L2. pH B.

ionize only

slightly in

solution £3.

titration C.

salt + H₂ (g) F

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4. Acids Bases

neutralization

D. $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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set (31) Aqueous
solutions acids
have a (blank)
taste. Sour
Taste. What
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Answers when

acids or bases
act as acids or
bases, they form
their conjugate
and the strength
is the inverse
of their parent.
and there are
certain strong
acids and strong

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bases. LiOH ,

NaOH , KOH – when
act as bases

(get rid of H^+
ion) then left
with Li Page

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Answers

4.8 as well

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Define acid and base according to Arrhenius's theory of ionization.

Explain the differences between strong and weak acids and bases.

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FIGURE 15-1 (a)
Fruits and fruit
juices contain
acids such as
citric acid and
ascorbic acid.
Carbonated
beverages
contain benzoic
acid, phosphoric
acid, and
carbonic acid.

(b) Many
household

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cleaners contain

•••
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Acids and Bases

Acids and Bases

Arrhenius

Definitions:

acids -

compounds that

produce an

increase in $[H^+]$

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when dissolved
in water bases -
compounds that
produce an

increase in
[OH⁻] when
dissolved in
water Lewis

Definitions:

acids - electron
pair acceptors

bases - electron
pair donors

Brønsted-Lowry

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Definitions:

acids – H⁺
donors

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Acids and Bases

a. Acid Base

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Answer Section

MULTIPLE CHOICE

1. ANS: B DIF: I

OBJ: 15-1.1 2.

ANS: C DIF: I

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OBJ: 15-1.1 3.

ANS: A DIF: I

OBJ: 15-1.1

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a. a strong acid

b. a weak acid

c. a strong base

d. a weak base

e. neither an

acid nor a base

24) A 25.0-mL

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sample of a solution of a monoprotic acid is titrated with a 0.115 M NaOH solution. The titration curve above was obtained. The concentration of the monoprotic acid is about _____ mol/L.

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